

The Crystal Structure of Riboflavin Hydrobromide Monohydrate

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The crystal structure of riboflavin hydrobromide monohydrate has been determined by a three-dimensional X-ray analysis. The space group is $P2_12_12_1$ with four formula units in the unit cell of dimensions $a=18.28$, $b=6.99$ and $c=14.77$ Å. The structural parameters were refined by means of a block-diagonal least-squares method, the final R factor being 0.07. The ribitol does not form a furanose ring but a chain. The protonation of the isoalloxazine ring occurs at the N(1) atom, and the C(4a)–N(10) bond is found to be a double bond. The bromide ion is situated midway between two isoalloxazine rings separated by a period along the b axis.

Riboflavin, in the form of the mono- and dinucleotide, is a prosthetic group of several enzymes concerning biochemical oxidation-reduction. These enzymes act as hydrogen carriers from a substrate to an acceptor molecule, when riboflavin alternatively accepts and releases two hydrogen atoms.

The electronic structure of riboflavin has been investigated in order to elucidate the mechanism of the transfer of the hydrogen atom. Guzzo and Tollin¹⁾ have suggested from the study of EPR spectra that it is reduced into semiquinone by strong acid. On the other hand, Pullman and Pullman have proposed the keto form to the reduced riboflavin from the theoretical calculation.²⁾

The present study has been undertaken to establish the relationship between the structural characteristics of mono-protonated riboflavin and its biological function.

Experimental

The crystals were obtained from 6N HBr solution of riboflavin. They are yellow rods elongated along the b axis. Unit cell dimensions were determined from the equatorial Weissenberg photographs about the b and c axes, on which the powder pattern of aluminum were superposed for calibration. The crystal data are shown in Table 1.

The equi-inclination Weissenberg photographs were taken for layer lines 0 to 5 about the b axis and 0 to 10 about c with $\text{CuK}\alpha$ radiation. The number of independent reflections recorded was 2261, but the intensities of 225 of these were too weak to be measured. The

TABLE 1. CRYSTAL DATA

Unit cell dimensions	$a=18.28\pm0.04$ Å $b=6.99\pm0.02$ $c=14.77\pm0.03$
Absent spectra	$(h00)$; $h=2n+1$ $(0k0)$; $k=2n+1$ $(00l)$; $l=2n+1$
Space group	$P2_12_12_1$
Density	$D_m=1.68$ g/cm ³ $D_c=1.67$ g/cm ³ , when $Z=4$
Molecular weight	475.3
Linear absorption coefficient for $\text{CuK}\alpha$ radiation	30.4
$F(000)$	976

relative intensities were measured visually and the range was 1 to 4000. They were corrected for Lorentz and polarization factors and spot-size factors for the equi-inclination Weissenberg photographs, but the corrections for absorption and secondary extinction were neglected. A rod-shaped crystal having a cross section of 0.5 mm \times 0.3 mm was used for the b axis photographs. A crystal for the c axis photographs, cut perpendicularly to the b axis, has a dimension of 0.2 mm \times 0.2 mm \times 0.3 mm.

Structure Determination and Refinement

From a three-dimensional Patterson function the bromide ion was easily located at $x=0.205$, $y=0.250$ and $z=0.119$. Fourier synthesis with the phases based on the bromide ion has an additional mirror plane because the y parameter of the heavy atom was a special value. The atoms in the isoalloxazine ring and C(1') in the ribityl group were located near the planes of $y=0.25$, while the as-

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1) A. V. Guzzo and G. Tollin, *Arch. Biochem. Biophys.*, **103**, 231 (1963).

2) B. Pullman and A. Pullman, *Proc. Natl. Acad. Sci. U. S.*, **45**, 136 (1959).

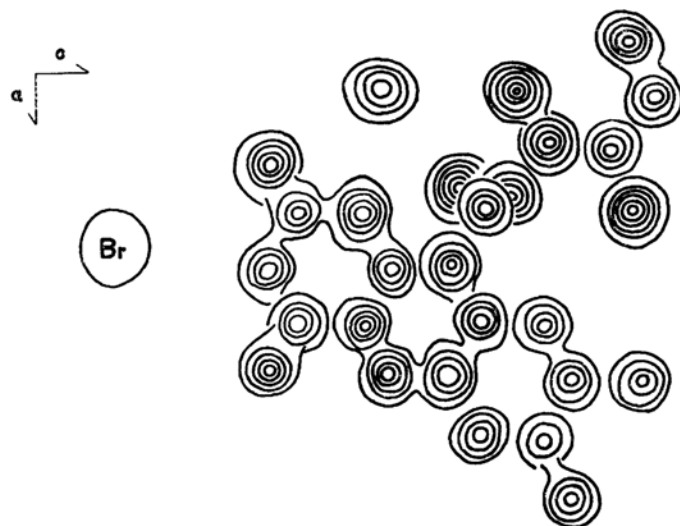


Fig. 1. Composite drawing of the final electron density map viewed along the b axis. Contours are drawn at intervals of $2.0 \text{ e} \cdot \text{\AA}^{-3}$; those for the bromide ion are not drawn. The lowest is $1.5 \text{ e} \cdot \text{\AA}^{-3}$.

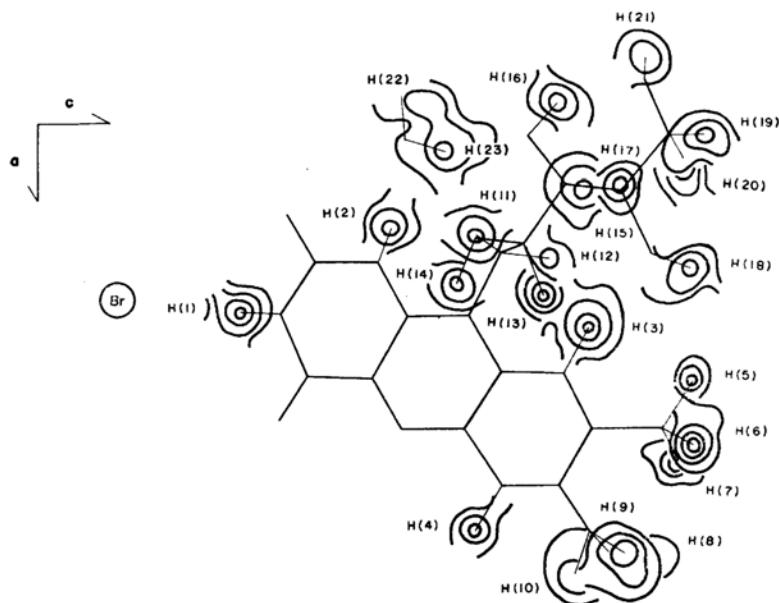


Fig. 2. Location of hydrogen atoms from a three-dimensional difference Fourier map. First contour $0.1 \text{ e} \cdot \text{\AA}^{-3}$; interval $0.2 \text{ e} \cdot \text{\AA}^{-3}$.

signment of the peaks to the other light atoms was somewhat puzzling because of the presence of the additional mirror plane. Successive Fourier summations were able to give well resolved peaks for all twenty-nine non-hydrogen atoms.

Atomic positions and anisotropic temperature factors were refined by a block-diagonal least-squares method. A difference Fourier map, calculated after several cycles of least-squares refinement, gave a fair indication of the positions of the hydrogen atoms. Further refinement, includ-

ing hydrogen atoms, was proceeded with anisotropic temperature factors for all the non-hydrogen atoms and with isotropic ones for hydrogen atoms. Atomic scattering factors were taken from the "International Tables for X-Ray Crystallography." Anomalous scattering effect was corrected for the bromide ion ($\Delta f' = -0.9$, $\Delta f'' = 1.4$). Weighting scheme was as follows;

$$\begin{aligned} w &= 1.0 & \text{if } |F_o| > 0 \\ w &= 0.5 & \text{if } |F_o| = 0 \end{aligned}$$

TABLE 2(a). ATOMIC PARAMETERS AND THEIR ESTIMATED STANDARD DEVIATIONS

Atom	x	y	z	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}
Br	0.2062(1)	0.2367(2)	0.1173(1)	234(3)	1960(26)	365(5)	382(21)	-52(8)	-11(28)
C(2)	0.1642(6)	0.2414(20)	0.3988(9)	212(31)	1705(239)	348(54)	2(190)	-51(70)	162(266)
C(4)	0.2953(6)	0.3106(15)	0.3945(8)	171(29)	939(197)	291(50)	152(144)	113(76)	122(179)
C(5)	0.4261(6)	0.2966(19)	0.6746(9)	179(32)	1111(305)	334(51)	5(170)	105(69)	-344(222)
C(6)	0.4294(6)	0.2904(19)	0.7664(9)	187(32)	1362(288)	337(51)	-16(173)	-2(70)	-5(226)
C(7)	0.3617(6)	0.2770(19)	0.8158(8)	210(33)	1205(249)	279(45)	35(188)	67(66)	-24(225)
C(8)	0.2964(6)	0.2714(17)	0.7740(7)	170(27)	1467(229)	273(42)	-170(197)	33(63)	209(212)
C(4a)	0.2966(6)	0.2916(15)	0.4950(7)	177(29)	979(231)	236(41)	-82(161)	47(65)	97(172)
C(8a)	0.2937(6)	0.2781(15)	0.6795(7)	175(27)	874(198)	232(39)	85(169)	6(62)	-299(178)
C(9a)	0.2287(5)	0.2651(17)	0.5400(8)	181(28)	1026(195)	277(43)	35(153)	7(57)	47(214)
C(10a)	0.3592(5)	0.2911(16)	0.6287(9)	140(27)	709(233)	370(52)	-111(141)	62(67)	174(221)
C(6m)	0.5009(6)	0.3012(24)	0.8182(9)	203(33)	2929(446)	334(55)	-198(214)	88(75)	25(283)
C(7m)	0.3632(6)	0.2697(20)	0.9200(8)	243(35)	1600(255)	293(47)	-268(208)	0(69)	-116(246)
N(1)	0.1659(5)	0.2370(16)	0.4921(6)	197(25)	1481(192)	238(36)	-125(149)	-35(51)	211(183)
N(3)	0.2281(5)	0.2744(15)	0.3533(6)	198(27)	1764(203)	241(36)	99(140)	25(51)	-92(173)
N(9)	0.2278(4)	0.2663(13)	0.6305(6)	142(21)	841(152)	284(37)	227(114)	46(47)	324(177)
N(10)	0.3584(5)	0.3028(14)	0.5360(6)	163(26)	972(230)	229(37)	-214(133)	58(54)	-105(161)
O(2)	0.1081(4)	0.2118(19)	0.3582(6)	206(24)	3889(361)	324(39)	-259(179)	-151(54)	552(234)
O(4)	0.3480(5)	0.3588(14)	0.3498(6)	224(26)	1872(227)	322(37)	-33(139)	167(51)	52(165)
C(1')	0.1573(5)	0.2585(19)	0.6838(7)	116(24)	1729(220)	257(41)	436(167)	20(54)	-375(228)
C(2')	0.1439(6)	0.0554(17)	0.7179(8)	135(29)	1411(254)	254(48)	158(152)	5(64)	-156(203)
C(3')	0.0775(6)	0.0442(17)	0.7774(8)	143(31)	1452(242)	285(49)	3(153)	29(67)	80(203)
C(4')	0.0844(6)	0.1629(18)	0.8607(9)	194(33)	1696(280)	303(55)	-116(169)	134(72)	4(209)
C(5')	0.0242(7)	0.1130(21)	0.9305(9)	250(40)	1857(327)	396(64)	25(207)	161(86)	376(263)
O(2')	0.1347(4)	-0.0664(12)	0.6397(6)	187(23)	1359(173)	311(38)	45(113)	-4(49)	-241(143)
O(3')	0.0168(4)	0.1002(14)	0.7258(6)	171(25)	2155(231)	277(37)	-6(135)	44(50)	96(173)
O(4')	0.1545(5)	0.1279(15)	0.9028(6)	204(24)	3052(276)	343(43)	102(150)	-18(56)	-200(193)
O(5')	-0.0462(4)	0.1137(13)	0.8916(8)	213(26)	1878(207)	588(55)	235(129)	245(70)	-53(217)
OW	0.0149(5)	0.1111(16)	0.5370(7)	240(30)	2700(277)	353(41)	113(161)	-49(61)	92(197)

Anisotropic temperature factors in the expression, $\exp\{-(h^2\beta_{11}+k^2\beta_{22}+l^2\beta_{33}+hk\beta_{12}+hl\beta_{13}+kl\beta_{23})\}$, are multiplied by 10^3 . The estimated standard deviations given in parentheses refer to the last decimal positions at respective values.

TABLE 2(b). THE FINAL PARAMETERS OF HYDROGEN ATOMS

Atom	**	x	y	z	Atom	**	x	y	z
H(1)	N(3)	0.2173	0.2480	0.2791	H(13)	C(2')	0.1886	0.0001	0.7500
H(2)	N(1)	0.1158	0.1801	0.5204	H(14)	O(2')	0.1840	-0.1250	0.6200
H(3)	C(8)	0.2422	0.2827	0.8004	H(15)	C(3')	0.0858	-0.0581	0.7855
H(4)	C(5)	0.4729	0.3214	0.6363	H(16)	O(3')	-0.0216	0.1003	0.7556
H(5)	C(7m)	0.3013	0.2675	0.9622	H(17)	C(4')	0.0808	0.2741	0.8508
H(6)	C(7m)	0.3752	0.3626	0.9610	H(18)	O(4')	0.1475	0.1694	0.9846
H(7)	C(7m)	0.3992	0.1286	0.9410	H(19)	C(5')	0.0439	0.1447	0.9967
H(8)	C(6m)	0.5145	0.1450	0.8551	H(20)	C(5')	0.0496	-0.0077	0.9629
H(9)	C(6m)	0.5069	0.4366	0.8548	H(21)	C(5')	-0.0652	0.1770	0.8858
H(10)	C(6m)	0.5324	0.3291	0.7826	H(22)	OW	-0.0285	0.0707	0.5273
H(11)	C(1')	0.1140	0.3097	0.6552	H(23)	OW	0.0258	0.0774	0.5868
H(12)	C(1')	0.1672	0.3218	0.7495					

* The mean value of the e.s.d. of the parameters are $\sigma(x)=0.13\text{ \AA}$, $\sigma(y)=0.15\text{ \AA}$ and $\sigma(z)=0.14\text{ \AA}$.

** The heavy atoms bonded to each hydrogen atom.

[illegible]

[illegible]

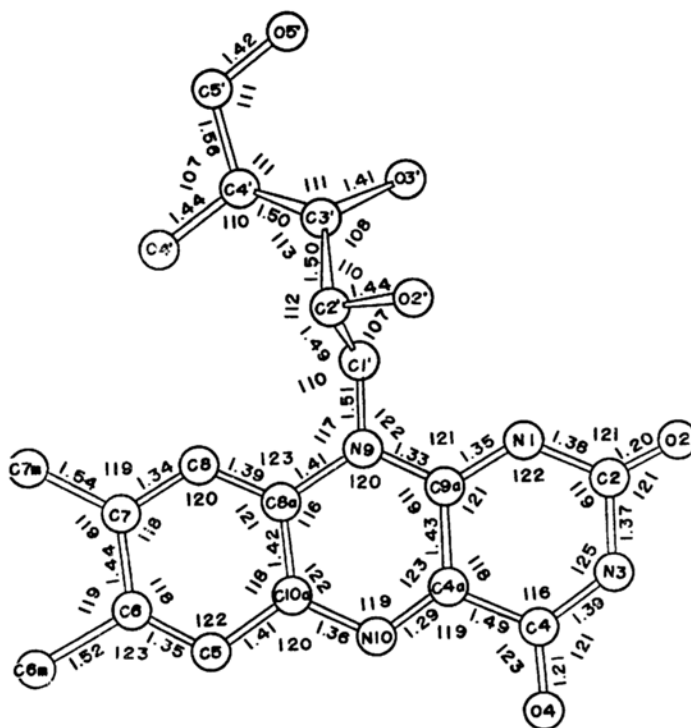


Fig. 3. Bond lengths (Å) and angles (degree).

the R factor was reduced to 0.07 after the refinement. Final atomic parameters are listed in Table 2, and $|F_o|$ and $|F_c|$ are tabulated in Table 3.

Discussion

The bond lengths and angles are shown in Fig. 3. The mean value of the estimated standard deviations for the bond lengths between non-hydrogen atoms except the bromide ion is 0.02 Å and that for the bond angles is 1.0°.

The difference Fourier synthesis (Fig. 2) shows that the addition of proton does not occur at the O(2) atom but at N(1) by hydrogen bromide. This is also supported by the fact that both exocyclic C-O lengths correspond to that of the double bond. The present result agrees with the location of the proton suggested by theoretical calculation,^{2,3)} and suggests that the pathway 1 in the reduction of riboflavin shown in Fig. 4 is reasonable, contrary to Guzzo and Tollin's expectation.¹⁾

The C(4)–C(4a) bond in the isalloxazine ring is 1.49 Å, which corresponds to the single bond length, and the C(4a)–N(10) bond is shorter than the mean value of the other C–N lengths. Therefore, it is presumed that π -electrons are localized at the C(4a)–N(10) bond.

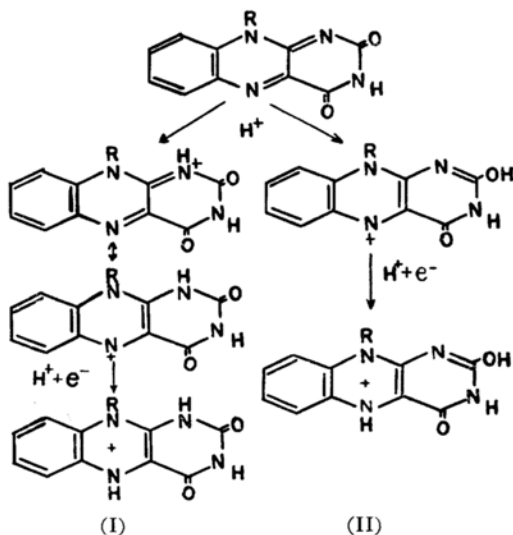


Fig. 4. Alternative pathways for the reduction of isoalloxazine.

On the theoretical calculation of the bond orders in the ring, the Hückel approximation was applied to a system of 16 conjugated π -electrons and 4 mobile electrons due to the hyperconjugation of the two methyl groups, where the Coulomb and exchange integrals were taken from the Ladik's⁴⁰. There are several relations proposed

3) B. Grabe, *Biopolymer symposia*, **No. 1**, 289 (1964).

TABLE 4. BOND ORDERS AND BOND LENGTHS

Bond	Bond order	Bond length	
		Theoretical	Experimental
N(1)—C(2)	1.34	1.38 Å	1.38 Å
N(1)—C(9a)	1.39	1.38	1.35
N(3)—C(2)	1.43	1.36	1.37
N(3)—C(4)	1.45	1.35	1.39
N(9)—C(8a)	1.54	1.33	1.41
N(9)—C(9a)	1.65	1.31	1.33
N(10)—C(4a)	1.70	1.30	1.29
N(10)—C(10a)	1.56	1.33	1.36
C(4a)—C(9a)	1.52	1.42	1.43
C(8a)—C(10a)	1.50	1.42	1.42
C(4)—C(4a)	1.30	1.47	1.49
C(5)—C(6)	1.73	1.38	1.35
C(5)—C(10a)	1.55	1.41	1.41
C(6)—C(7)	1.59	1.40	1.44
C(7)—C(8)	1.72	1.38	1.34
C(8)—C(8a)	1.56	1.41	1.39

for converting the bond order into length.⁵⁻¹⁰⁾ For C—C bonds, approximately the same values were obtained by all the relations, and the converted bond lengths by, say, Coulson's⁵⁾ are in good agreement with those observed, as shown in Table 4. On the other hand, the converted C—N lengths show somewhat wider variations with the relations, and agreement between the theoretical and observed lengths is not so good. But if experimental errors are taken into account, significant disagreement is only for N(9)—C(8a). This may suggest that this approximation is not necessarily satisfactory in the present molecule.

The valence angles at the N(1) and N(3) atoms with an extra-annular hydrogen atom are 125° and 122°, respectively, and that at N(10) without hydrogen is 119°. They are consistent with the nitrogen valence angle in heterocyclic compounds discussed by Singh.¹¹⁾ There are some small valence angles at the C(8a) and C(4) atoms which are bonded to the nitrogen atoms with an extra-annular hydrogen atom, whereas slightly large ones are found at the C(4a) and C(10a) atoms

bonded to the N(10) atom. These facts have been reported in alloxane,¹²⁾ anhydrous alloxane,¹³⁾ anhydrous barbituric acid,¹⁴⁾ barbituric acid dihydrate,¹⁵⁾ pyridimine,¹⁶⁾ and 2-amino-4,6-dichloropyridine.¹⁷⁾

The best planes of the isoalloxazine ring and the six-membered rings are listed in Table 5 together with the deviations of the atoms from them, where the nomenclatures of the rings are as follows:

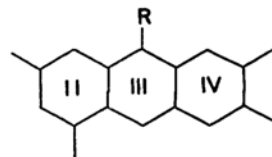


TABLE 5. THE BEST PLANE OF THE RINGS

Atom	Plane I	Plane II	Plane III	Plane IV
N(1)	0.097	-0.014	0.153*	
C(2)	0.119	0.014		
N(3)	0.039	0.019		
C(4)	-0.109	-0.038	-0.100*	
C(4a)	-0.035	0.037	-0.015	
N(10)	-0.020	0.135*	-0.012	-0.033*
C(10a)	0.007		0.025	-0.001
C(5)	0.068		0.073*	0.001
C(6)	0.063			0.001
C(7)	-0.002			0.001
C(8)	-0.062		-0.011*	-0.001
C(8a)	-0.057		-0.016	0.000
N(9)	-0.071	-0.089*	-0.017	0.045*
C(9a)	-0.007	-0.025	0.037	
C(6m)	0.092*			
C(7m)	-0.012*			
O(2)	0.242*			
O(4)	-0.315*			
C(1')	-0.185*			

$$\text{Plane I}^{**} \quad 0.104X - 0.994Y - 0.041Z = -1.717$$

$$\text{Plane II} \quad 0.175X - 0.984Y - 0.040Z = -1.375$$

$$\text{Plane III} \quad 0.089X - 0.995Y - 0.033Z = -1.765$$

$$\text{Plane IV} \quad 0.053X - 0.998Y - 0.036Z = -2.010$$

Dihedral angles between the planes

$$\text{Plane I—Plane II} \quad 4.1 \text{ (degree)}$$

$$\text{Plane I—Plane III} \quad 1.0$$

$$\text{Plane I—Plane IV} \quad 3.0$$

$$\text{Plane II—Plane III} \quad 5.0$$

$$\text{Plane II—Plane IV} \quad 7.1$$

$$\text{Plane III—Plane IV} \quad 2.1$$

* Not included in the calculation of the equation of the plane.

** X, Y and Z in the Å unit.

12) C. Singh, *ibid.*, **19**, 759 (1965).

13) W. Bolton, *ibid.*, **17**, 147 (1964).

14) W. Bolton, *ibid.*, **16**, 166 (1963).

15) G. A. Jeffrey, S. Ghose and J. O. Warwicker, *ibid.*, **14**, 881 (1961).

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8) T. H. Goodwin and A. L. Porte, *J. Chem. Soc.*, **1956**, 3595.

9) T. Anno, M. Ito, R. Shimada, A. Sado and W. Mizushima, *This Bulletin*, **30**, 638 (1957).

10) M. J. S. Dewar and H. N. Schmeising, *Tetrahedron letters*, **5**, 166 (1959).

11) C. Singh, *Acta Cryst.*, **19**, 861 (1965).

and the I represents the isoalloxazine ring. The dihedral angles between the planes are also given in this table. Though the ring IV is strictly planar, the other two rings are puckered; the deviations of

the atoms from their best planes are significant. The dihedral angles between the planes indicate that the isoalloxazine ring takes, though slightly, a boat form. However, the isoalloxazine ring in

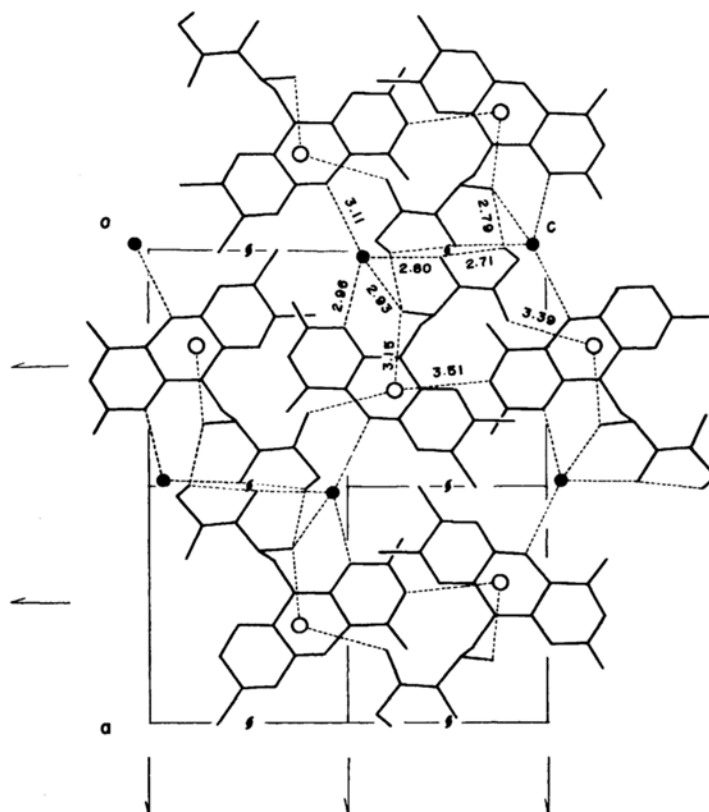


Fig. 5. The packing of the molecules viewed along the *b* axis. The dotted lines show the possible hydrogen bonds.

Medium circle: bromide ion, black circle: water molecule

TABLE 6. LIST OF POSSIBLE HYDROGEN BONDS

X—HY	X—Y	X—H	∠XHY
Br(i)—H(14, v)O(2', v)	3.15 Å	2.52 Å	151 degree
Br(i)—H(1, i)N(3, i)	3.51	2.40	171
Br(i)—H(18, ii)O(4', ii)	3.39	2.28	146
OW(i)—H(2, i)N(1, i)	2.96	1.92	162
O(2', i)—H(23, i)OW(i)	2.93	2.36	130*
O(3', i)—H(23, i)OW(i)	2.80	2.06	151*
N(10, iii)—H(22, i)OW(i)	3.11	2.43	162
O(5', i)—H(16, i)O(3', i)	2.71	2.06	135**
O(2, iv)—H(21, i)O(5', i)	2.79	2.22	177

Key for molecules

i: *x*, *y*, *z*

ii: *x*, *y*, 1+*z*

iii: $-1/2+x$, $1/2-y$, $1-z$

iv: $-x$, $1/2+y$, $3/2-z$

v: $1/2-x$, $-y$, $-1/2+z$

* May be a bifurcated hydrogen bond

** The intramolecular hydrogen bond

1,3,10-trimethyl isalloxazine iodide¹⁸⁾ has been reported to be nearly coplanar.

In the ribityl group, the C(1')-C(2'), C(2')-C(3') and C(3')-C(4') bonds are a little shorter than the normal C-C length of sp^3 hybridization, whereas the C(4')-C(5') does not differ significantly

from the normal one. The C-O lengths in this group are between 1.41 Å and 1.44 Å, and their mean value, 1.428 Å, corresponds to the usual single C-O length. The valence angles C-C-C and C-C-O in this group do not differ significantly from their mean value, 107° and 110°, respectively.

The ribitol does not form a furanose ring but a chain, which has not been found in cases of nucleoside, nucleotide and nucleic acid. The conformation angles defined by Sundaralingam¹⁹⁾ are; $\phi_{C_2'-C_3'} = 62^\circ$ and $\phi_{C_3'-C_4'} = 167^\circ$, whereas they are 34° and -37° in the furanose ring of 5'-AMP.²⁰⁾

The molecular packing viewed along the *b* axis is shown in Fig. 5. The hydrogen bonds are shown by the dotted lines, where each figure represents a distance between the non-hydrogen atoms. Table 6 tabulates the X-H distances and X-H-Y angles related to the hydrogen bond. The H(2,i)-OW(i), 1.92 Å, is the normal hydrogen bond. The H(23,i) is located near both O(2',i) and O(3',i) atoms; the distances are 2.36 Å and 2.06 Å, respectively. Thus, the two contacts, OW-H(23)-O(2') and OW-H(23)-O(3'), seem to consist a bifurcated hydrogen bond. Another hydrogen atom of water molecule makes a weak hydrogen bond with the N(10) atom related by the screw axis along the *a*. This hydrogen bond may suggest the subsequent transfer of the hydrogen atom from the substrate.

The bromide ion contributes to the three hydrogen bonds with the N(3), O(2') and O(4') atoms. Though the Br(i)-N(3,i) distance, 3.51 Å, seems a little longer, the Br(i)-H(1,i) is 2.40 Å, which shows that it is a normal hydrogen bond. The bromide ion is situated midway between two isalloxazine rings separated by a period along the *b* axis, and nearly above the center of the ring III which is indicated to have positive charge from the theoretical calculation of the bond orders.

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TABLE 7. CLOSE INTERMOLECULAR CONTACTS

Atom	Nearest atom	Molecular position	Distance
Br	C(6m)	iii	3.87
O(2)	C(5)	iii	3.35
O(2)	C(6m)	iii	3.26
O(3')	O(4)	iii	3.29
OW	O(4)	iii	3.48
C(1')	O(5')	iv	3.39
Br	C(8a)	v	3.70
Br	C(9a)	v	3.86
Br	C(10a)	v	3.87
Br	N(9)	v	3.71
Br	C(2')	v	3.72
C(4)	O(4')	v	3.19
C(4a)	O(4')	v	3.34
O(2)	C(7)	v	3.50
O(4')	C(2')	v	3.48
O(4)	C(3')	v	3.30
O(4)	O(4')	v	3.48
Br	C(4a)	vi	3.75
Br	C(8a)	vi	3.50
Br	C(9a)	vi	3.84
Br	C(10a)	vi	3.50
Br	N(9)	vi	3.67
Br	N(10)	vi	3.62
C(2)	C(7m)	vi	3.45
N(3)	C(8)	vi	3.40
O(5')	C(7m)	vii	3.34

Key for molecules

iii: $-1/2+x, 1/2-y, 1-z$

iv: $-x, 1/2+y, 3/2-z$

v: $1/2-x, -y, -1/2+z$

vi: $1/2-x, 1-y, -1/2+z$

vii: $-1/2+x, 1/2-y, 2-z$

This table shows the contacts shorter than 4.0 Å between the bromide ion and the other heavy atom, and those shorter than 3.5 Å in the other cases.

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